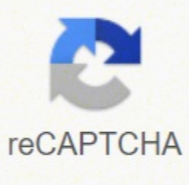




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Mole Conversions Worksheet #1

1. Mole -> Mass Conversions – using molar mass of each substance, convert the following quantities.

- | | |
|---|--|
| a. 10.0 mol Cr 520 g | f. 0.160 mol H ₂ O 2.88 g |
| b. 3.32 mol K 130 g | g. 5.08 mol Cu(NO ₃) ₂ 834 g |
| c. 2.20×10^{23} mol Sn 0.261 g | h. 15.0 mol H ₂ SO ₄ 1470 g |
| d. 0.720 mol Be 6.48 g | i. 4.52×10^5 mol C ₂ H ₄ 1.27×10^3 g |
| e. 2.40 mol N ₂ 67.2 g | j. 0.0112 mol K ₂ CO ₃ 1.55 g |

2. Mass -> Mole Conversions – using molar mass of each substance convert the following quantities.

- | | |
|---|--|
| a. 72.0 g Ar 1.80 mol | f. 27.4 g NO ₂ 0.596 mol |
| b. 3.70×10^{23} g B 3.43×10^{-2} mol | g. 5.00 g H ₂ 2.50 mol |
| c. 187 g Al 6.93 mol | h. 2.64×10^{-2} g Li ₃ PO ₄ 2.28×10^{-4} mol |
| d. 333 g SnF ₂ 2.13 mol | i. 11.0 g CH ₄ 0.688 mol |
| e. 7.21×10^{23} g He 1.80×10^{-2} mol | j. 847 g (NH ₄) ₂ CO ₃ 8.82 mol |

3. What is the volume of the following gases?

- 5.40 mol O₂ 121 L
- 3.20×10^{23} mol CO₂ 0.717 L
- 0.960 mol SO₂ 21.5 L

4. How many moles are in each of the following volumes?

- 89.6 L Ne 4.00 mol
- 1.00×10^3 L C₂H₄ 44.6 mol
- 5.42×10^3 L F₂ 2.42×10^3 mol

5. Find the number of moles in each of the number of representative particles.

- 1.20×10^{23} atoms of P 19.9 mol
- 3.87×10^{23} molecules of AlF₃ 6.43×10^3 mol
- 4.81×10^{23} molecules of NH₃ 7.99×10^{10} mol

6. How many representative particles are in each of the following mole quantities?

- 1.24 mol Cl₂ 7.46×10^{23} molecules
- 4.20×10^{23} mol K₂S 2.53×10^{23} molecules
- 34.02 mol Ca(OH)₂ 2.048×10^{25} molecules

7. Convert the following two-step quantities, converting first to moles and then to the desired quantity.

- Find the number of molecules in 60.0 g of N₂O. 8.21×10^{23} molecules
- Find the volume of 3.24×10^{23} molecules of Ne 1.21 L
- Find the mass of 18.0 L of CH₄ 12.9 g
- Find the volume of 835 g of SO₂ 234 L
- Find the mass of one atom of nickel. 1×10^{22} g

Name: _____
Date: _____

Stoichiometry Worksheet #1 Answers

1. Given the following equation: $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$, show what the following molar ratios should be.

- C₄H₁₀ / O₂
- O₂ / CO₂
- O₂ / H₂O
- C₄H₁₀ / CO₂
- C₄H₁₀ / H₂O

2. Given the following equation: $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$

- How many moles of O₂ can be produced by letting 12.00 moles of KClO₃ react?
18.0 mol O₂

3. Given the following equation: $2 \text{K} + \text{Cl}_2 \rightarrow 2 \text{KCl}$

- How many grams of KCl is produced from 2.50 g of K and excess Cl₂?
4.77 g KCl
- How many grams of KCl is produced from 1.00 g of Cl₂ and excess K?
2.10 g KCl

4. Given the following equation: $\text{Na}_2\text{O} + \text{H}_2\text{O} \rightarrow 2 \text{NaOH}$

- How many grams of NaOH is produced from 1.20×10^3 grams of Na₂O?
154.8 g NaOH
- How many grams of Na₂O are required to produce 1.60×10^3 grams of NaOH?
124 g Na₂O

5. Given the following equation: $8 \text{Fe} + \text{S}_8 \rightarrow 8 \text{FeS}$

- What mass of iron is needed to react with 16.0 grams of sulfur?
27.87 g Fe
- How many grams of FeS are produced?
43.9 g FeS

Question: How exactly does a buffer work? How can its pH not increase when a strong base (i.e. NaOH) is added?

Answer: The net result is that free OH- ions are not allowed to accumulate. For example, OH-(aq) in the last sample problem is not around for that long, but instead it gets replaced by the much weaker base, C2H3O2-, according to:



because such a strong base.

The Henderson-Hasselbalch Equation.

In a buffered solution, the pH is governed by the ratio of [HA] (weak acid) to the common ion, [A-].

pH depends on the ratio [HA]/[A-].

for problem-solving, this is integrated into the Henderson-Hasselbalch Equation:

pH = pKa + log([A-]/[HA])

this is a quick and easy way to replace ICE Method for buffers.

ex: Using the Henderson-Hasselbalch equation, calculate the pH of a buffer solution containing 0.50M HC2H3O2 and 0.50M NaC2H3O2. Ka for acetic acid = 1.8 x 10^-5.

pH = pKa + log([C2H3O2-]/[HC2H3O2])

pH = 4.74 + log(1)

pH = 4.74 matches ICE Method result on p. 15-3.

pKa = -log(Ka)
pKa = -log(1.8 x 10^-5)
pKa = 4.74

- 1. What is the mass of 7.5 moles of sulfur dioxide (SO2)?
7.5 mol * 64g = 480g
2. How many grams of sodium phosphate (Na3PO4) are there in 1.52 mol?
250g / 164g = 1.52 mol
3. How many grams of potassium sulfate (K2SO4) are there in 2.5 moles?
25.3 mol * 176g = 4402g
4. What is the volume of 0.25 moles of gas at STP?
0.25 mol * 22.4L = 5.6L
5. What is the mass of 0.1 moles of copper?
0.1 mol * 63.5g = 6.35g
6. Calculate the number of moles in 32.4 g of SO2.
32.4g / 64g = 0.506 mol
7. Calculate the number of grams in 0.25 mol of AgNO3.

ChemistryNotes.com

Molar Ratio Worksheet

Find the molar ratio for the following compounds:

- 1. HCl
2. FeCl3
3. Hg2Cl2
4. BaPO3
5. (NH4)2CO3
6. K2SO4
7. Fe(OH)3
8. NaOH
9. NaCl
10. Ba(OH)2

Molar ratio practice problems worksheet answers.

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